

Chapter 13 States Of Matter

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Chapter 13 States Of Matter

Chapter 13 States of Matter 137 SECTION 13.1 THE NATURE OF GASES (pages 385–389) This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance. Kinetic Theory and a Model for Gases (pages 385–386) 1.

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all matter consists of tiny particles that are constantly in motion What are the three assumptions of the kinetic theory as it applies to gases? -The particles in a gas are considered to be small, hard spheres with an insignificant volume. -The motion of the particles in a gas are rapid, constant, and random.

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There are three states of matter that we will learn about in this chapter. (If you want to learn about more states of matter, I can refer you to somebody.) Those three states are solid, liquid, and gas. These three states are quite different. The main difference is in their particles.

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No attractive or repulsive forces exist between the particles. 3 Chapter 13 States Of Matter Worksheet Title: Chapter 13 States of Matter 1 Chapter 13 States of Matter 2 Kinetic Theory as Applied to Gases Fundamental assumptions about gases. The particles in a gas are considered to be small, hard spheres with an insignificant volume.

Chapter 13 States Of Matter Worksheet

Title: Chapter 13 States of Matter 1 Chapter 13 States of Matter 2 Kinetic Theory as Applied to Gases Fundamental assumptions about gases. The particles in a gas are considered to be small, hard spheres with an insignificant volume. Between particles in a gas there is empty space. No attractive or repulsive forces exist between the particles. 3

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Chapter 13 States of Matter. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. bill_brossman. Terms in this set (26) Kinetic Energy. The energy an object has because of its motion. Kinetic Theory. A theory that explains the states of matter, based on the concept that all matter consists of tiny particles that are ...

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Chapter 13 - States of Matter - 13.3 The Nature of Solids - 13.3 Lesson Check - Page 434: 24 Answer Solids have a definite shape and a definite volume because of the arrangement of the particles.

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Chapter 13 - States of Matter - 13 Assessment - Page 443: 54 Answer Temperature stays constant because the energy is needed to break the intermolecular forces and changing the phase.

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Title: Chapter 13: States of Matter 1 Chapter 13 States of Matter.
Kinetic-Molecular Theory Explains the motions and behavior of a
gas. The theory has three components ; 1. Particle Size Gas
particles are small relative to the space around the particles.
This means that there is no significant attraction or repulsion
between gas particles. NO ...

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Section 13.1 continued CHAPTER 13 STUDY GUIDE FOR CONTENT
MASTERY Vacuum Atmospheric pressure Pressure exerted by
mercury column 760 mm Name Date Class Study Guide for
Content Mastery Chemistry: Matter and Change Chapter 13 73
States of Matter Section 13.1 Gases In your textbook, read about
the kinetic-molecular theory. Complete each statement. 1.

13 STUDY GUIDE FOR CONTENT MASTERY 13 STUDY GUIDE FOR ...

Chapter 13 "States of Matter". Use these activities to help
yourself learn the vocabulary and major concepts presented in
this chapter. graph representing the relationships among the
solid, liquid, and vapor states of a substance in a sealed
container. This activity was created by a Quia Web subscriber.

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